

# 100% Every Student Every Day

|   | Monday  | Tuesday  | Wednesday  |
|---|---|--|--|
| <b>TEKS Dual Coding</b>                             | <b>SE:</b> 8.5(C) interpret the arrangement of the Periodic Table, including groups and periods, to explain how properties are used to classify elements<br><b>(D)</b> recognize that chemical formulas are used to identify substances and determine the number of atoms of each element in chemical formulas containing subscripts. | <b>SE:</b> 8.5(D) recognize that chemical formulas are used to identify substances and determine the number of atoms of each element in chemical formulas containing subscripts;<br><b>(F)</b> recognize whether a chemical equation containing coefficients is balanced or not and how that relates to the law of conservation of mass. | <b>SE:</b> 8.5(D) recognize that chemical formulas are used to identify substances and determine the number of atoms of each element in chemical formulas containing subscripts;<br><b>(F)</b> recognize whether a chemical equation containing coefficients is balanced or not and how that relates to the law of conservation of mass. |
|   | <b>Process Standard 8.3(B)</b>  | <b>Process Standard 8.3(B)</b>   | <b>Process Standard 8.3(D)</b>   |
| <b>Lesson Objective (WE will learn)</b>             | We will learn to recognize that chemical formulas are used to identify substances and determine the number of atoms of each element in chemical formulas containing subscripts.   | We will learn to recognize that chemical formulas are used to identify substances and determine the number of atoms of each element in chemical formulas containing subscripts.  | We will learn recognize whether a chemical equation containing coefficients is balanced or not and how that relates to the law of conservation of mass.  |
| <b>I will statement (Demonstration of learning)</b> | I will finish Cornell notes and complete 8.5D, F STEMscopes Part I: The Big Picture Student Guide and Student Journal.  | I will complete 8.5D, F STEMscopes Part II: Chemical Formulas, and Part III: Chemical Equations Student Guide and Student Journal.   | I will take the <b>Interactive Notebook (INB) Quiz</b> and Complete Reading assignment The Birth of Chemical Formulas and Equations  |
| <b>Purposeful Instructional Agenda</b>              | <ol style="list-style-type: none"> <li>1. Warm up</li> <li>2. Finish Cornell Notes</li> <li>1. 8.5D, F STEMscopes Part I: The Big Picture Student Guide and Student Journal</li> </ol>  | <ol style="list-style-type: none"> <li>1. Warm up</li> <li>2. 8.5D, F STEMscopes Part II: Chemical Formulas, and Part III: Chemical Equations Student Guide and Student Journal</li> </ol>   | <ol style="list-style-type: none"> <li>1. <b>INB Quiz on DMAC – BRING YOUR DEVICE!!!</b></li> <li>2. Reading assignment The Birth of Chemical Formulas and Equations</li> </ol>  |
|   | <b>Homework: Read STEMscopedia due Friday</b>   | <b>Homework: Prepare for Interactive Notebook (INB) Quiz</b>   | <b>Homework: None</b>  |
| <b>Seed Question FSGPT</b>                          | What is a subscript and what does it indicate in a chemical formula?  | What is the law of conservation of mass and how does it relate to chemical equations?  | How are coefficients used in a chemical equation?  |
| <b>AVID strategy</b>                                | <b>Cornell Notes</b>  |  | <b>Reading to Learn</b>  |
| <b>Kagan Strategy</b>                               | <b>Partner Share 6 Word Summary</b>   |  | <b>Independent</b>   |

## 100% Every Student Every Day

|   | Thursday   | Friday   | Notes                              |
|---|--|--|------------------------------------|
| <b>TEKS<br/>Dual Coding</b>                             | <p><b>SE:</b> 8.5(D) recognize that chemical formulas are used to identify substances and determine the number of atoms of each element in chemical formulas containing subscripts;<br/>                     (F) recognize whether a chemical equation containing coefficients is balanced or not and how that relates to the law of conservation of mass.</p> <p><b>Process Standard 8.3(B)</b></p> | <p><b>SE:</b> 8.5(D) recognize that chemical formulas are used to identify substances and determine the number of atoms of each element in chemical formulas containing subscripts;<br/>                     (F) recognize whether a chemical equation containing coefficients is balanced or not and how that relates to the law of conservation of mass.</p> <p><b>Process Standard 8.3(B)</b></p> | <b>Fundraiser kick-off Tuesday</b> |
| <b>Lesson Objective<br/>(WE will)</b>                   | We will learn to recognize whether a chemical equation containing coefficients is balanced or not and how that relates to the law of conservation of mass  | We will learn to recognize whether a chemical equation containing coefficients is balanced or not and how that relates to the law of conservation of mass  |                                    |
| <b>I will statement<br/>(Demonstration of learning)</b> | I will work on 8.5D, F STEMscopes Part III: Chemical Equations Student Guide and Student Journal.  | I will complete 8.5D, F STEMscopes Part IV: The Balancing Act Student Guide and Student Journal, and Reflections and Conclusions.  |                                    |
| <b>Purposeful Instructional Agenda</b>                  | <ol style="list-style-type: none"> <li>1. Warm up</li> <li>2. 8.5D, F STEMscopes Part III: Chemical Equations Student Guide and Student Journal</li> </ol>   | <ol style="list-style-type: none"> <li>1. Warm up</li> <li>2. Complete 8.5D, F STEMscopes Part IV: The Balancing Act Student Guide and Student Journal</li> <li>3. Reflections and Conclusions</li> </ol>  |                                    |
| <b>Homework:</b>  | <b>None</b>  | <b>STEMscopedia <u>DUE</u></b>   |                                    |
| <b>Seed Question<br/>FSGPT</b>                          | Why is it necessary to balance a chemical equation?  | What role do coefficients play in balancing a chemical equation?   |                                    |
| <b>Avid Strategy</b>                                    | <b>Cornell Notes</b>   | <b>Collaboration – complete reflections as a group</b>   |                                    |
| <b>Kagan Strategy</b>                                   | <b>Partner Share 6 Word Summary</b>  | <b>Shoulder Partners-Face Partners</b>   |                                    |