1. ***Atom*** - The smallest unit of an element that has all of the properties of that element containing a nucleus within an electron cloud.
2. ***Atomic Mass*** - The average mass of protons and neutrons in an element.
3. ***Atomic Number*** - The number of protons in the nucleus of one atom of an element.
4. ***Chemical Symbol*** - A one or two letter notation used to represent an atom of a particular element.
5. ***Electrical Charge*** - A form of charge, designated negative, positive, or neutral (without charge) that is found on the subatomic particles that make up all atoms.
6. ***Electron*** - A negatively charged subatomic particle of the electron cloud that is involved in the formation of chemical bonds.
7. ***Electron Cloud/Energy Level/Electron Shell*** - The regions around the nucleus where electrons may be found.
8. Elem***ent*** - A pure substance that cannot be separated into simpler substances by physical or chemical means.
9. ***Groups*** - The columns on a Periodic Table that arrange the elements by the number of electrons that are in the outermost shell.
10. ***Ions -*** Form as a result of the loss or gain of electrons; identified by the overall net charge.
11. ***Metalloids -*** Elements that have properties of both metals and non-metals; sometimes referred to as semiconductors.
12. ***Metals -*** Most elements are metals; they are typically solid, shiny, malleable, and good conductors of heat and electricity.
13. ***Net Charge*** - The sum of negative and positive charges.
14. ***Neutron*** - A subatomic particle of the nucleus of an atom that is without charge that contributes to the mass of an atom.
15. ***Noble Gases*** - Unreactive non-metals in Group 18 of the Period Table.
16. ***Non-Metals*** - Elements typically not shiny, usually a gas or brittle solid, not malleable, and poor conductors of heat and electricity.
17. ***Nucleus*** - The tiny, very dense, positively charged region in the center of an atom; made up of protons and neutrons.
18. ***Periodic Table of Elements*** - A table showing the chemical elements arranged according to their atomic numbers.
19. ***Periods*** - The rows in a Periodic Table that classify the elements by the number of electron shells.
20. ***Proton*** - A positively charged subatomic particle of the nucleus of an atom that contributes to the mass of the atom.
21. ***Reactivity*** - Rate at which a chemical substance tends to undergo a chemical reaction; significantly influenced by valence electrons of the reacting substances.
22. ***Subatomic Particles*** - Particles that are smaller than the atom.
23. ***Valence Electron Shell*** - The partially- filled outermost shell (or shells) determine the chemical properties of the atom.
24. ***Valence Electrons*** - The electrons in the outermost energy level of an atom that influence how an element will react with other substances.